

**Table 9.3****Common Polyatomic Ions**

<b>Formula</b>	<b>Name</b>
<b>Charge = 1<sup>-</sup></b>	
$\text{H}_2\text{PO}_4^-$	Dihydrogen phosphate
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate
$\text{HSO}_3^-$	Hydrogen sulfite
$\text{HSO}_4^-$	Hydrogen sulfate
$\text{HCO}_3^-$	Hydrogen carbonate
$\text{NO}_2^-$	Nitrite
$\text{NO}_3^-$	Nitrate
$\text{CN}^-$	Cyanide
$\text{OH}^-$	Hydroxide
$\text{MnO}_4^-$	Permanganate
$\text{ClO}^-$	Hypochlorite
$\text{ClO}_2^-$	Chlorite
$\text{ClO}_3^-$	Chlorate
$\text{ClO}_4^-$	Perchlorate
<b>Charge = 2<sup>-</sup></b>	
$\text{HPO}_4^{2-}$	Hydrogen phosphate
$\text{C}_2\text{O}_4^{2-}$	Oxalate
$\text{SO}_3^{2-}$	Sulfite
$\text{SO}_4^{2-}$	Sulfate
$\text{CO}_3^{2-}$	Carbonate
$\text{CrO}_4^{2-}$	Chromate
$\text{Cr}_2\text{O}_7^{2-}$	Dichromate
$\text{SiO}_3^{2-}$	Silicate
<b>Charge = 3<sup>-</sup></b>	
$\text{PO}_3^{3-}$	Phosphite
$\text{PO}_4^{3-}$	Phosphate
<b>Charge = 1<sup>+</sup></b>	
$\text{NH}_4^+$	Ammonium