

**Table 11.3****Solubility Rules for Ionic Compounds**

<b>Compounds</b>	<b>Solubility</b>
Salts of alkali metals and ammonia	Soluble
Nitrate salts and chlorate salts	Soluble
Sulfate salts, except compounds with $\text{Pb}^{2+}$ , $\text{Ag}^+$ , $\text{Hg}_2^{2+}$ , $\text{Ba}^{2+}$ , $\text{Sr}^{2+}$ , and $\text{Ca}^{2+}$	Soluble
Chloride salts, except compound with $\text{Ag}^+$ , $\text{Pb}^{2+}$ , and $\text{Hg}_2^{2+}$	Soluble
Carbonates, phosphates, chromates, sulfides, and hydroxides	Most are insoluble