

Table 8.3**Electronegativity Differences and Bond Types**

| Electronegativity difference range | Most probable type of bond | Example |
|---|-----------------------------------|--|
| 0.0–0.4 | Nonpolar covalent | H—H (0.0) |
| 0.4–1.0 | Moderately polar covalent | $\overset{\delta+}{\text{H}} - \overset{\delta-}{\text{Cl}}$ (0.9) |
| 1.0–2.0 | Very polar covalent | $\overset{\delta+}{\text{H}} - \overset{\delta-}{\text{F}}$ (1.9) |
| ≥ 2.0 | Ionic | Na^+Cl^- (2.1) |