- Atomic Number : # of Protons
  : # of electrons
- Mass Number (whole #) : # Protons + Neutrons
- Atomic Mass (decimal) : Weighted Average of mass of Isotopes
- Isotope : element w/ different # of neutrons

^{11}Na 22.98

^{23}Na

^{11}

Sodium-23

Short Answers

① X-35 X-40

Two isotopes

Atomic Mass = 39,42 close

Which is more abundant

X-40

② Element X 10% ^{40}X, 20% X^{42}

70% ^{45}X

Find Atomic Mass:

$.10(40) + .20(42) + .70(45) =$

% mass

③ Given # of neutrons and the mass #

Find atomic #

Atomic # = Mass # − neutrons

Essay

In what way are two isotopes of the same element different? Explain why isotopes of the same element have the same chemical behavior.